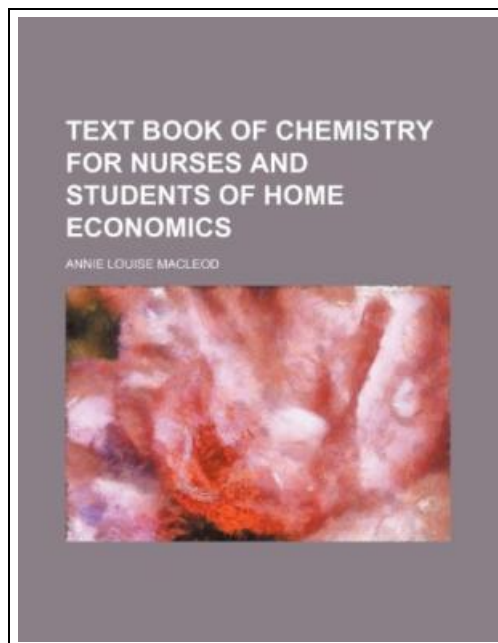


Text book of chemistry for nurses and students of home economics



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Reviews

The ebook is easy in read through preferable to understand. It is actually writter in straightforward words and never hard to understand. I realized this publication from my dad and i encouraged this ebook to understand.
(Dr. Fausto Jenkins Sr.)

TEXT BOOK OF CHEMISTRY FOR NURSES AND STUDENTS OF HOME ECONOMICS



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RareBooksClub. Paperback. Book Condition: New. This item is printed on demand. Paperback. 46 pages. Dimensions: 8.9in. x 6.0in. x 0.6in. This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1920 edition. Excerpt: . . . acid produced in this way, containing 3-6 per cent, of acid together with flavoring material from the fruit or grain extract used for fermenting. The name vinegar, without qualification, is used in the United States for the product made from apple juice or cider only, but vinegar may also be made from beer (malt vinegar), wine (white wine vinegar), or molasses. Acetic acid is also obtained, along with methyl alcohol, when wood is destructively distilled, and it is in this way that most of the acetic acid used industrially is obtained. The acid produced when milk sours is lactic acid, $\text{CH}_3\text{CHOH}\cdot\text{COOH}$. This acid has a hydroxyl group in the molecule as well as a carboxyl, and therefore shows some of the properties of an alcohol as well as those of an acid. Tartaric acid, present in the form of its sodium-potassium salt in grape juice, from which it is extracted and sold as cream-of-tartar, contains two hydroxyl and two carboxyl groups, $\text{COOH}\cdot\text{CHOH}\cdot\text{CHOH}\cdot\text{COOH}$. Citric acid, to which oranges and lemons owe their sour taste, has one hydroxyl and three carboxyl groups. Organic acids are neutralized by inorganic bases, forming salts, $\text{CH}_3\cdot\text{COOH}\cdot\text{NaOH}\cdot\text{CH}_3\text{COONa}\cdot\text{H}_2\text{O}$. A dicarboxylic acid, such as tartaric, takes two molecules of base for neutralization; a tricarboxylic acid takes three molecules of base. $\text{COOH}\cdot\text{CHOH}\cdot\text{CHOH}\cdot\text{COOH}\cdot 2\text{NaOH}\cdot\text{COONa}\cdot\text{CHOH}\cdot\text{CHOH}\cdot\text{COONa}\cdot\text{H}_2\text{O}$. Here the hydroxyl, being alcoholic in its nature, is unacted upon by the base....



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